

Name: \_\_\_\_\_

## Notes on Acids and Bases SNC2D

Acids:

- taste \_\_\_\_\_
- are \_\_\_\_\_  
(react with \_\_\_\_\_ to produce \_\_\_\_\_)
- release \_\_\_\_\_ ions (\_\_\_\_\_) in solution

Examples of acids:

\_\_\_\_\_

Formula: \_\_\_\_\_

\_\_\_\_\_

Formula: \_\_\_\_\_

\_\_\_\_\_

Formula: \_\_\_\_\_

\_\_\_\_\_

Formula: \_\_\_\_\_

Naming acids:

**Binary acids** (hydrogen + a single element) are called hydro\_\_\_\_\_ic acids

e.g.  $\text{H}_2\text{S}_{(\text{aq})}$  is \_\_\_\_\_ acid

**Oxyacids** (hydrogen + a polyatomic ion) are called:

\_\_\_\_\_ic acids if the name of the ion ends in -ate

\_\_\_\_\_ous acids if the name of the ion ends in -ite

e.g.  $\text{HNO}_{3(\text{aq})}$  is \_\_\_\_\_ acid

$\text{HNO}_{2(\text{aq})}$  is \_\_\_\_\_ acid

Bases:

- taste \_\_\_\_\_
- feel \_\_\_\_\_
- are \_\_\_\_\_ (break down \_\_\_\_\_)
- release \_\_\_\_\_ ions (\_\_\_\_\_) in solution

Examples of bases:

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The strength of an acid or base may be measured on the **pH scale**.



The scale runs from \_\_\_\_\_ to \_\_\_\_\_. Neutral solutions (e.g. water) have a pH of \_\_\_\_\_.

Acidic solutions have a pH of \_\_\_\_\_.

The closer the pH to 7, the \_\_\_\_\_ the acid.

A basic or \_\_\_\_\_ solution has a pH \_\_\_\_\_.

The closer the pH to 7, the \_\_\_\_\_ the base.

It is possible to test the pH of a solution using indicators, e.g. \_\_\_\_\_.

Blue litmus paper turns \_\_\_\_\_ in an acid and remains blue in neutral or alkaline solutions.

Red litmus paper turns \_\_\_\_\_ in a base and remains red in neutral or acidic solutions.